

PART I- TRANSPORT PROCESSES: MOMENTUM, HEAT, AND MASS

(1.2-1)

Eq. (1.3-4)  $K = 353.2 = ^\circ C + 273.15$   $^\circ C = 80.05 = \boxed{80^\circ C}$

Eq. (1.3-2)  $^\circ C = 80.0 = (1/1.8)(^\circ F - 32)$   $^\circ F = \boxed{176^\circ F}$

Eq. (1.3-3)  $^\circ R = ^\circ F + 460 = 176 + 460$   $^\circ R = \boxed{636^\circ R}$

(1.2-2)

Eq. (1.3-2)  $^\circ C = (1/1.8)(^\circ F - 32) = (1/1.8)(155 - 32) = \boxed{68.33^\circ C}$

Eq. (1.3-4)  $K = ^\circ C + 273.15 = 68.33 + 273.15 = \boxed{341.5 K}$

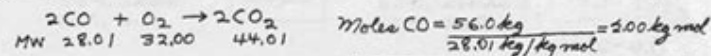
Eq. (1.3-3)  $^\circ R = ^\circ F + 460 = 155 + 460 = \boxed{615^\circ R}$

(1.3-1)

$MW(O_2) = 32.00, MW(N_2) = 28.02$

$MW(\text{air}) = 0.21(32.00) + 0.79(28.02) = \boxed{28.9 \text{ kg/kg mol}}$

(1.3-2)



$Moles \ O_2 = 1.00, \quad 1.00(32.00) = \boxed{32.0 \text{ kg } O_2}$        $2.00(44.0) = \boxed{88.0 \text{ kg } CO_2}$

(1.3-3)

Gas	g	MW	g mol	Mol frac
N <sub>2</sub>	20	28.02	0.7138	0.1648 mol frac
O <sub>2</sub>	83	32.00	2.5938	0.5990
CO <sub>2</sub>	45	44.01	1.0225	0.2362
Total	148 g		4.3301 g mol	1.0000

Average mol wt =  $\frac{148 \text{ g}}{4.3301 \text{ g mol}} = \boxed{34.2 \text{ g/g mol}}$   
 $\boxed{34.2 \text{ kg/kg mol}}$

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