

## Chemistry Semester 1 Final Exam Review

- What is a definition of Chemistry?  
**Study of matter**
  - Matter includes all of the following EXCEPT:  
a. air      **b. light**      c. smoke      d. water vapor
  - What is matter?  
**Something that has mass and takes up space (volume)**
  - A physical change occurs when a  
a. peach spoils      c. copper bowl tarnishes  
b. bracelet turns your wrist green      **d. glue gun melts a glue stick**
  - The particles in a solid are  
**a. packed closely together**      c. constantly in motion  
b. very far apart      d. able to slide past one another
- Of the following groups, 1, 2, 14, 15, 17, 18:
- Group(s) **1, 2** contains only metals
  - Group (s) **18** is called a Noble gases
  - Group(s) **17** has 7 electrons on the valence shell
  - Group(s) **15** needs to gain 3 electrons to fill a stable octet
  - Group(s) **2** has a 2+ oxidation state
  - Group(s) **2** has an electron dot notation of  $\bullet X \bullet$
  - Using the periodic table, an element similar to carbon would be **14**
  - Based on their location on the periodic table, you could infer that **18/noble gases** are very unreactive
  - The horizontal row on the periodic table is called a(n) **period**
  - The vertical column of blocks on the periodic table is called a(n) **group/family**
  - The elements on the zigzag line in the periodic table are **metalloids**
  - Which is NOT a property of metal?  
a. malleability      **c. unreactivity**  
b. ability to lose electrons      d. ability to conduct heat
  - A volume of 1 cc ( $\text{cm}^3$ ) is the same as **1** mL
  - To determine density, the quantities that must be measured are **volume** and **mass**
  - The relationship between the mass  $m$  of a material, its volume  $V$ , and its density  $D$  is  
a.  $V = mD$       b.  $Vm = D$       **c.  $DV = m$**       d.  $D + V = m$
  - The density of aluminum is 2.70  $\text{g/cm}^3$ . The volume of a solid piece of Aluminum is 1.50  $\text{cm}^3$ . Find the mass **4.05 grams**
  - The density of pure diamond is 3.5  $\text{g/cm}^3$ . The mass of a diamond is 0.25 g. Find the volume= **0.71  $\text{cm}^3$**
  - What statement about density is true?  
a. two samples of a pure substance may have different densities if they are different sizes  
b. the density of a sample depends on its location on Earth  
c. a cylinder is always used to measure the volume  
**d. density is a physical property that remains constant for a pure substance**
  - The number of grams equal to 0.5 kg is **500 g**
  - In a graph the **independent** variable is placed on the x axis and the **dependant** variable is placed on the y axis.
  - The variable that you change in the lab is the **independent** variable. The variable that responds to the change is called the **dependant** variable. All other things do not change and are called **constants**
  - At first Dalton thought the atom was solid, spherical, **indivisible**, then Rutherford discovered the **nucleus** of the atom, then the subatomic particles **protons**, **neutrons**, and **electrons** were discovered
  - The positively charged particle is the **proton** found in the **nucleus** of the atom
  - The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a **neutron** and is found in the **nucleus** of the atom
  - The **electron** is negatively charged and is found in the **electron cloud** outer shell of the atom
  - Bhor** is the person credited with placing electrons on levels and **Schrodinger** established the electron cloud theory.
  - The forces that hold the particles together in the nucleus are **columbic /electrostatic** forces
  - Isotopes are atoms of the same element that **have different number of neutrons**
  - All isotopes of hydrogen contain **1 proton & 1 electron**
  - What is the atomic number for Nitrogen? **7** To what subatomic particle does this correspond? **# of protons**
  - Zinc -66 has **30** protons and **36** neutrons
  - What is the electron configuration for sodium?  **$[\text{Ne}]3s^1$**  Chlorine?  **$[\text{Ne}] 3s^2 3p^4$**  Neon?  **$[\text{He}] 2s^2 2p^6$**
  - What is the dot notation for magnesium? Bromine? Argon?

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# Semester Exam Review Answer Key For Chemistry

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### **Ninth grade Lesson Semester Review Questions | BetterLesson**

Answer Key A.P. Statistics Semester I Review 1 In the paper "Reproduction in Laboratory colonies of Bank Vole," the authors presented the results of a study of litter size. (A vole is a small rodent with a stout body, blunt nose, and short ears.) As each new litter was born, the number of babies was recorded, and the accompanying results were obtained.

### **Answer Key A.P. Statistics Semester I Review**

Chemistry Semester 1 Exam Review Answer Key Italics indicate required honors extensions Essential Questions are answered within the review by performing the practice Honors chemistry semester 1 exam review answer key. Vocab is found in the text. Honors chemistry semester 1 exam review answer key

### **Honors Chemistry Semester 1 Exam Review Answers**

GEOMETRY B Semester Exam Review Key © MCPS 1. a. same volume b. same volume 2. 25 78.54 3. 112 351.86 4. a. circle b. triangle c. rectangle d. square

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### **ALG 2B Algebra II, Second Semester To the Student: WHAT TO ...**

About once every 6 months new questions are added to the unit tests and we update this section first. Some of the lecture answer key pairs include: Polynomials, Factoring, Relations and Matrices.

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2014-2015 Answer Key: Answer Key to English III Review for Semester 1. A student's final grade in the course will be calculated using the EOCT as follows (State Board Rule 160-4-2-. Bienvenidos a mi página web. If you were absent, make sure you see me to see our class set of notes so you can ask questions as they ...

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