

9. If the pH of a sulfuric acid solution is 1.2,
What is the concentration of hydronium?

$$[\text{H}_3\text{O}^+] = 10^{-1.2} = .063 \text{ M}$$

What is the hydroxide concentration?

$$[\text{OH}^-] = 10^{-12.8} = 1.58 \times 10^{-13} \text{ M}$$

What is the molarity of the solution?

$$\frac{.063}{2} = .0315 \text{ M}$$

What is the pOH?

$$14 - 1.2 = 12.8$$

10. If the pH of a solution of sodium hydroxide is 9.0,
What is the concentration of hydronium?

$$[\text{H}_3\text{O}^+] = 10^{-9.0} = 1.0 \times 10^{-9} \text{ M}$$

What is the hydroxide concentration?

$$[\text{OH}^-] = 10^{-5.0} = 1.0 \times 10^{-5} \text{ M}$$

What is the molarity of the solution?

$$1.0 \times 10^{-5} \text{ M}$$

What is the pOH?

$$14 - 9.0 = 5.0$$

11. If the pH of a solution of ammonia is 7.6,
What is the concentration of hydronium?

$$[\text{H}_3\text{O}^+] = 10^{-7.6} = 2.5 \times 10^{-8} \text{ M}$$

What is the hydroxide concentration?

$$[\text{OH}^-] = 10^{-6.4} = 4.0 \times 10^{-7} \text{ M}$$

What is the molarity of the solution?

$$4.0 \times 10^{-7} \text{ M}$$

What is the pOH?

$$14 - 7.6 = 6.4$$

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Mole calculations - Formula mass and mole calculations ...

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Molecules, Moles and Avogadro's Number

Mole, standard unit (6.02214076 x 10²³) in chemistry for measuring large quantities of very small entities such as atoms, molecules, or other specified particles. The number of units in a mole also bears the name Avogadro's number, or Avogadro's constant, in honor of the Italian physicist Amedeo Avogadro.

mole | Definition, Number, & Facts | Britannica

Chemistry- Moles and Equations Test. moles. Avogadro. 6.02 x 10²³. use its mass number. in chemistry, everything is related to ____ Who calculated the number of atoms found in a mole? What is Avogadro's number? What do you use to find out how many grams are in 1 mole of an... moles. in chemistry, everything is related to ____

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Test Chemical Reactions And Balancing Equations

number of moles = c x v = 0.100 x 0.0274 = 0.00274 moles step 4: Using the ratio, how many moles of sulphuric acid for every 2 NaOH there are 1 H 2 SO 4 so, we must have 0.00274/2 =0.00137 moles of H 2 SO 4 Step 5: Calculate concentration. concentration = moles/volume in dm³ = 0.00137/0.025 = 0.0548 moldm⁻³

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