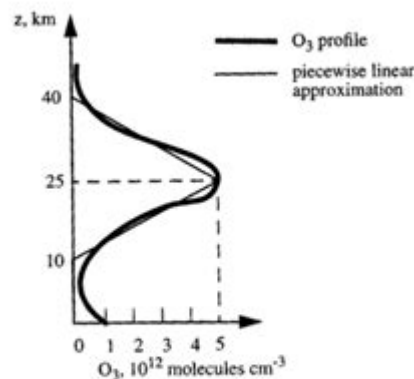


PROBLEM 1 - (from introduction to atmospheric chemistry by Daniel Jacob):

Consider the typical vertical profile of ozone number density measured over the US shown in the opposite figure. This layer protects life on Earth by absorbing solar UV radiation. Inhalation of ozone is toxic to humans and plants and therefore, there is a 65 ppb standard for ground-level ozone concentration in Canada.



a) Calculate the mixing ratio of ozone at the peak concentration altitude ($z=25$ km, $P=35$ hPa, $T = 220$ K). Would this mixing ratio be in violation of Canadian standards?

b) Calculate the mixing ratio of ozone in surface air ($z=0$ km, $P=1000$ hPa, $T=300$ K). Is it in compliance with Canadian standards? Notice that the relative decrease in mixing ratio between 25 km and the surface is considerably larger than the relative decrease in number density. Why is this?

c) The total number of ozone molecules per unit area is called ozone column and determines the efficiency with which the ozone layer prevents solar UV radiation from reaching the surface. Estimate the ozone column for the above profile by approximating the profile with the piecewise linear function shown as the thin solid line.

d) To illustrate how thin this stratospheric ozone layer actually is, imagine that all of the ozone in the atmospheric column were brought to sea level as a layer of pure ozone gas under standard conditions (1013 hPa and 273 K). Calculate the thickness of this layer.

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Daniel J. Jacob, Supplemental Problems for "Introduction to Atmospheric Chemistry", 5th edition, 2012. 1 INTRODUCTION TO ATMOSPHERIC CHEMISTRY: SUPPLEMENTAL QUESTIONS AND PROBLEMS 5th EDITION . by Daniel J. Jacob . Harvard University . August 2012 . FOREWORD . The questions and problems presented here are intended to supplement my book

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