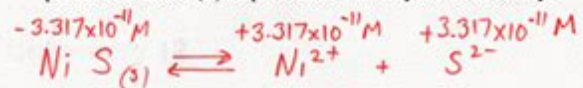


10. The molar solubility of nickel (II) sulphide is 3.317×10^{-11} M. Calculate the value of the solubility product for nickel (II) sulphide. Show all of your work clearly.



$$K_{sp} = [\text{Ni}^{2+}][\text{S}^{2-}]$$

$$K_{sp} = (3.317 \times 10^{-11})^2 = \boxed{1.100 \times 10^{-21}}$$

11. A solution of potassium chloride is titrated with 0.200 M silver nitrate solution. The following data table was obtained:

	Trial 1	Trial 2	Trial 3
Initial AgNO ₃ burette reading (mL)	0.00	5.26	14.63
Final AgNO ₃ burette reading (mL)	5.26	12.19	19.87
Volume of KCl titrated	25.0	25.0	25.0

Use the information in the data table to calculate the [Cl⁻] in the KCl solution. Show all of your work clearly.

	Trial 1	Trial 2	Trial 3
Vol. AgNO ₃ used (mL)	5.26	6.93 DISCARD	5.24

$$\text{Best average volume} = \frac{5.26 + 5.24}{2} = 5.25 \text{ mL}$$

$$= \underline{0.00525 \text{ L AgNO}_3}$$

$$\text{moles of AgNO}_3 = 0.200 \text{ M} \times 0.00525 \text{ L} = \underline{0.00105 \text{ mol AgNO}_3 (\text{Ag}^+)}$$



$$\text{moles KCl} = 0.00105 \text{ mol AgNO}_3 \times \frac{1 \text{ mol KCl}}{1 \text{ mol AgNO}_3}$$

$$= 0.00105 \text{ mol KCl}$$

$$[\text{Cl}^-] = [\text{KCl}] = \frac{0.00105 \text{ mol}}{0.0250 \text{ L}} = \boxed{0.0420 \text{ M}}$$

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Acid Base Titration Simulation Lab Answer Key

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Acid Base Titration Simulation Lab

2. Explain the term acid-base titration. 3. Write balanced chemical equations representing acid-base reactions. 4. Solve acid-base titration problems involving molarity, solution volume, and number of moles of solute (acid and base). 5. Calculate the concentration of a solute (acid or base) given information provided by a titration experiment.

Acid-Base Titration Computer Simulation | Chemdemos

you are here->home->Chemical Sciences->Inorganic Chemistry Virtual Lab->Acid Base Titration .. Acid Base Titration .. Theory . Procedure . Self Evaluation . Simulator . Assignment . Reference . Feedback ...

Acid Base Titration (Simulator) : Inorganic Chemistry ...

In chemistry laboratory, it is sometimes necessary to experimentally determine the concentration of an unknown acid or base solution. A procedure for making this kind of determination is called an acid-base titration.

Lab # Acid-Base Titration Simulation Introduction

In this simulation, you will embrace the power of the titration technique. With appropriate lab apparatus, a chemical indicator and a base solution, you can determine the concentration of any acid. Assemble the apparatus for titration A successful experiment starts with good preparation.

Virtual Lab: Titration Virtual Lab | Labster

Your instructor Pang, the chemistry cat, will help you through your first experiment: an acid/base titration. Simply click the "Run Game" button to get started. In simulation mode, you'll have full access to all of the laboratory equipment and chemicals you'll need to perform titrations of strong/weak acid and bases.

Virtual Titration by scientistsarepeopletoo

The titration screen experiment has been designed to be a free flexible tool for teachers and students. You can choose to carry out a strong acid - strong base titration (or any combination of strong and weak acid-base titrations). There is also a redox titration experiment to complete in order for students to practise their understanding and ...

Titration screen experiment | Resource | RSC Education

Analytical Chemistry/Lab Techniques. Acid/Base Titrations; Standardization of NaOH with a KHP solution: Acid Base Titration . Use the Virtual Laboratory to standardize an unknown NaOH solution (approximately 0.2M) to four significant figures via titration with 25.00 mL of a KHP standard solution.

ChemCollective: Virtual Labs

To begin, select the acid, the base, their concentrations, which of them is in the burette, and choose an indicator. Then click "Add solutions...". Click on the burette stopcock to start, pause or stop the titration. (The titration will stop automatically at the 25 mL mark if it is not stopped earlier.) Any parameter or parameters can be changed after

the first have been chosen. Click "Add ...

Titration Simulator - Concordia University

Given acids or bases at the same concentration, demonstrate understanding of acid and base strength by: 1.Relating the strength of an acid or base to the extent to which it dissociates in water 2.Identifying all of the molecules and ions that are present in a given acid or base solution. 3.Comparing the relative concentrations of molecules and ions in weak versus strong acid (or base ...

Acid-Base Solutions - Acids | Bases | Equilibrium - PhET ...

The chemical reaction involved in acid-base titration is known as neutralisation reaction. It involves the combination of H³O⁺ ions with OH⁻ ions to form water. In acid-base titrations, solutions of alkali are titrated against standard acid solutions. The estimation of an alkali solution using a standard acid solution is called acidimetry.

Acid Base Titration - Amrita Vishwa Vidyapeetham Virtual Lab

Recorded with <https://screencast-o-matic.com>

Virtual Lab Acid & Base Titration - Part 1 - YouTube

Users can model and simulate chemical reactions, focusing on thermodynamics, equilibrium, kinetics, and acid-base titrations, with accompanying virtual lab exercises. It is designed for high school (AP/IB) and undergraduate students and teachers.

Virtual Chemistry and Simulations - American Chemical Society

Acid-Base Titration: Interactive Simulation Name: ____ Date: ____ Discussion The molar concentration (Molarity) of an acid solution is determined by neutralization with a base of known concentration. A standard base solution is added to a known quantity of an acid solution until the reaction is complete, as shown by a sudden change in the color of an acid-base indicator. This sudden change is

Acid-Base Titration Simulation - irion-isd.org

acid_base - University of Oregon

acid_base - University of Oregon

Titration screen experiment - Royal Society of Chemistry

A titration is a process used to determine the volume of a solution that is needed to react with a given amount of another substance. In this experiment, your goal is to determine the molar concentration of two acid solutions by conducting titrations with a base of known concentration. You will be testing a strong acid, HCl, solution and a weak acid, HC₂H₃O₂, solution.

Acid-Base Titration - Vernier

Acid/ Base Titration Lab Design III Materials: Goggles HCl NaOH Distilled Water Label Beaker Graduated Cylinder Pipet Stirring rod Buret with clamp Stopper Ring Stand Utility Clamp Computer Computer Program pH probe Procedure: 1) 2) Goggles were

put on. 3) 0.1 M NaOH solution was made using 0.4 g NaOH and 100 mL distilled water. The ingredients were added together in a beaker and stirred ...

Lab Report Acid Base Titration Essay - 1352 Words

Titration involves an acid and a base, and the equivalence point is when the number of moles added in the base equals to the number of moles of acid in the titrant. To determine when the equivalence is reached an indicator is added to the titrant. Indicators are usually weak organic acids that have a different colour that its conjugate base.

Acid Base Titration Lab Report.docx - 1 Name Mahak Kapoor ...

Acid and Base Titrations Lab Report. University. University of Miami. Course. Chemistry Laboratory I (CHM 113) Uploaded by. Apples Bananas. Academic year . 2015/2016. Helpful? 88 77. Share. Comments. Please sign in or register to post comments. Related documents. A Green Chemistry Experiment Acid & base titration lab Chemical Kinetics - lab report Determination of Cu²⁺ - lab report ...

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